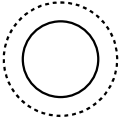
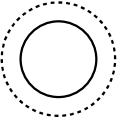
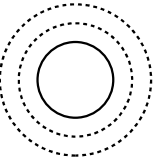
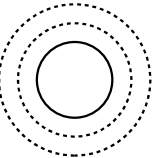
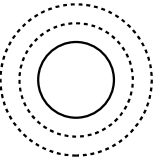
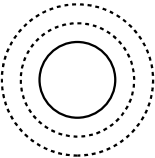
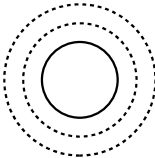
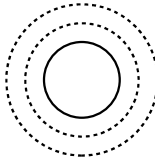
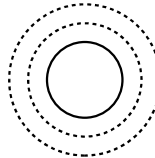
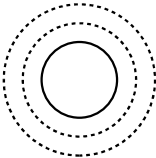
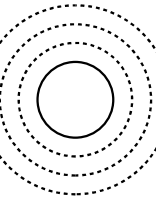
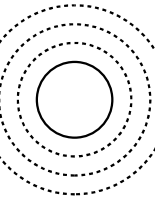
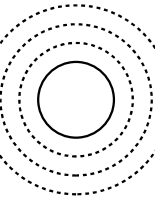
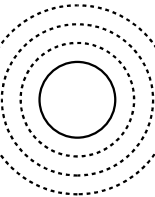
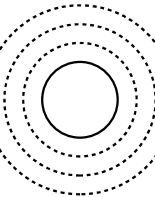
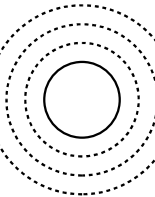
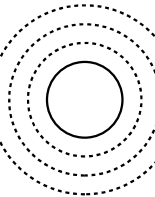
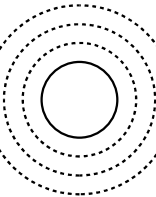
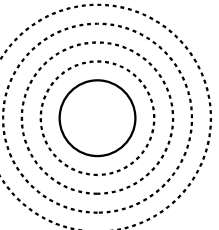
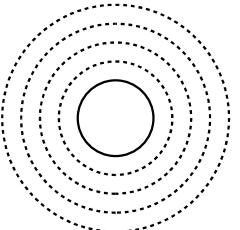


Orbital (ring) #1 holds 2 electrons

Orbital (ring) #2 holds 8 electrons

Orbital (ring) #3 holds 8 electrons

Orbital (ring) #4 holds 18 electrons

	Group 1								Group 18
Period 1									
Period 2									
Period 3									
Period 4									

On a separate piece of paper discuss the following periodic patterns:

- 1) Group number (vertical column) and number of valence electrons (outer shell electrons)
- 2) Period number (horizontal row) and the number of orbitals (shells)
- 3) The maximum number of electrons in each shell and the number of elements in each period.