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Bohr-Rutherford Diagrams

Date:		

Orbital (ring) #1 holds 2 electrons Orbital (ring) #2 holds 8 electrons Orbital (ring) #3 holds 8 electrons Orbital (ring) #4 holds 18 electrons

	Group 1							Group 18
Period 1		Group 2	Group 13	Group 14	Group 15	Group 16	Group 17	
		Group 2	Group 15	Group 11	Group 15	Group ro	Group 17	
Period 2								
Period 3								
Period 4								

On a separate piece of paper discuss the following periodic patterns:

- 1) Group number (vertical column) and number of valence electrons (outer shell electrons)
- 2) Period number (horizontal row) and the number of orbitals (shells)
- 3) The maximum number of electrons in each shell and the number of elements in each period.